

Original Research Paper

Sorption Extraction of Noble and Non-Ferrous Metals from Process Leaching Solutions

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Abstract: In view of the annual increase in ore processing output for low grade or refractory gold-bearing ores, as well as recoverable raw materials, stricter requirements for completeness of gold extraction and environmental safety, the interest is rising in possible alternatives to substitute the cyanide process used in gold mining practice. One of such alternatives is to convert metals into solutions of various chemical compositions by means of electric current-electrochemical leaching-followed by metal sorption extraction. This paper deals with the issues of sorption behavior of noble (gold, silver) and non-ferrous (copper, iron, cobalt, zinc, nickel, manganese, lead) metals from thiosulfate solutions for chemical and electrochemical leaching of concentrates of two mineral deposits using various anionites. Factors affecting metal uptake most significantly were determined. Degree of zinc and nickel sorption extraction from electrochemical leaching solutions made 99%. It has been established that the nature of metal sorption extraction from process solutions is connected with the micro structure of leaching solutions having quasi-homogeneous structure.

Keywords: Leaching Solutions, Noble Metals, Non-Ferrous Metals, Sorption

Introduction

The global trend in mineral resources base development lies in increasing the volume of raw mineral sources involved in processing generally by means of lower grade and compositionally complicated deposits. It is possible to increase mineral resources base of gold using low-grade, refractory and man-made raw materials, which may be achieved as a result of efficiency improvement of the processing technology and completeness of valuable component extraction. The most widespread ore refractory behavior factor is caused by gold availability inside sulfide mineral, primarily in pyrite and arsenopyrite, with minimum size of the embedded gold grains ranging from decimicron to centimicron. Among the scientific and technical problems facing modern gold mining industry, the problem of extracting gold, silver and non-ferrous metals from recoverable, man-made and refractory mineral stock may be referred to the most significant ones (Adams, 2005; Vignes, 2013). Wide occurrence of gold-bearing ores with ultradispersed inclusions of valuable components in the nature necessitates their large-scale involvement into processing. Complicated multi-stage

processing flow sheets for such types of refractory ores prevent from extracting ultradispersed gold completely; therefore in most cases such mineral stock shall remain unprocessable and refers to substandard ores according to the economic and environmental criteria. Commercial development of compositionally complicated ores with low content of gold (1.0-2.0 g/t) using conventional processing technology is accompanied with huge losses of valuable and associated components having grain size less than 40 μm .

The purpose of the present research is to study sorption of thiosulfate complexes of noble and non-ferrous metals from thiosulfate solutions for chemical and electrochemical leaching with various anion-exchange resins; to investigate sorption properties of a number of new anion-exchange resins.

Literature Review

Development of a number of technologies, including heap leaching, enables to involve ores containing about 1 g/t of gold in ore processing and achieve its extraction up to 60-80%. However, processing of gold-bearing ores of numerous deposits is restrained by refractory behavior

factor for the subsequent gold extraction by hydrometallurgical methods. Practically all methods of vein gold processing have considerable disadvantages resulting in technological, economic and environmental problems during their implementation. Complicated processing flow sheets are energy intensive, mainly due to multi-stage additional grinding of products. Presence of arsenopyrite in the majority of ores limits the possibility to process them according to gravity separation and flotation schemes because of strict requirements for flotation concentrate with regard to arsenic content and practically excludes pyrometallurgy.

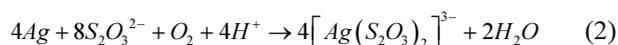
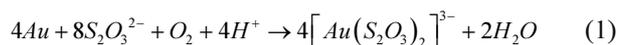
One of the main ways to increase process efficiency of the noble and non-ferrous metals extraction from refractory ore is improvement of modern hydrometallurgical processes with appropriate efficient methods of metal extraction from solutions. Currently the primary way to build up competitiveness and profitability of noble and non-ferrous metals production is to combine conventional technology with innovative ones: Heap and electrosorption leaching, electrochemical extraction, hydrometallurgical solution processing, etc. In addition, due to increasing range of application of non-ferrous, rare and precious metals stricter requirements are imposed on their purity. Production of pure metals is complicated by the fact that because of the high-grade ore lack metallurgical recovery practice involves natural raw materials containing negligible amount of valuable components (1-3 g per 1 t of ore and even less) (Lorenzen and Van Deventer, 2003). In many cases the existing pyro- and hydrometallurgical methods fail to ensure the required purity of the product extracted from such raw materials, as well as comprehensive extraction of valuable substances from polymetallic ores and concentrates (Bringas and Wayman, 1998; Su, 2012).

Cyanide process widely used both in the domestic and foreign industry is the basis of modern metallurgy of gold (Syed, 2012). Environmental aspects arising in connection with the use of cyanide solvents related to highly toxic substances are becoming increasingly pressing. In addition to the environmental hazard cyanide leaching method has a number of other shortcomings. Thus, to prevent cyanide ion hydrolysis the presence of protective alkalis is required, i.e., maintenance of a high pH-value of solutions, which makes it impossible to apply efficient oxidizers and therefore the process of gold dissolution has a longer duration (above 72 h). High tendency of cyanide ion to form complexes with ions of non-ferrous metals leads to additional consumption of reagent and during gold sorption concentrating makes it difficult to selectively remove it from the solution and complicates the regeneration of ion-exchange materials, leading to the loss of the precious metal with the "tailings". This results in the need to seek new ways of leaching processes using less environmentally hazardous solvents (Kyle *et al.*,

2012). Gold is often associated with sulfides and requires using flotation followed by autoclave or bacterial leaching of concentrates (Xia, 2008). Sulfide ores are difficult to process by conventional methods. According to the research, the intensity of leaching copper and zinc from sulfide ores is commensurable with this factor for mixed and oxidized ores (0.01-0.5% of copper per day and 0.1-1.0% of zinc per day, respectively) (Harvey *et al.*, 1992; Lalancette *et al.*, 1999; Conić *et al.*, 2014).

In view of the above it becomes necessary to carry out integrated processing of refractory sulfide gold concentrates and to substitute cyanide with less expensive and less toxic thiosulfate solvent, intensifying the process rate and reducing costs for detoxification of refined products. Hydrometallurgical preparation of sulfide concentrates allows obtaining chemical products, from which gold can be extracted by means of leaching with non-cyanide complex-forming reagents. A number of reagents were offered as a possible alternative to cyanide, including thiosulfates, thiourea (Yuce, 1994), halides (Tran and Davis, 1992), polysulfides (Luo *et al.*, 2006) and thiocyanates (Akcil *et al.*, 2015). The greatest interest in the global practice of gold extraction from natural materials is generated by the leaching process where ammonium or sodium thiosulfate is used as the main reagent (Grosse *et al.*, 2003; Zhang and Dreisinger, 2003; Petersen and Muzawazi, 2014; Lampinen *et al.*, 2015), since precisely this method is acquiring practical importance for the conditions of the in-situ leaching. Thiosulfate leaching possesses obvious advantages, both in terms of technological and environmental parameters (Wan, 1997; West-Sells and Hackl, 2005).

The thiosulfate leaching process is based on oxidation of gold and silver with atmospheric oxygen over thiosulfate in acidic medium:



Accordingly, there is a number of patents on thiosulfate leaching of crude ore (Wan *et al.*, 1995; Choi *et al.*, 2009). Depending on the pH of the medium, thiosulfate may be unstable and converted to polythionates. US Patent of 1976 is interesting in this regard, where, in contrast to many others, alkaline ammonia solution rather than acid one was offered for ore uncovering (Dubeck and Im, 1976).

Formed thiosulfate gold complex is very strong (dissociation constant is equal to 10^{-26}). The presence of soluble copper and sulfides can slow down the process of thiosulfate dissolution of gold, if special measures are not taken. In particular, it is recommended to carry out the process in a weak oxidizing medium.

Thiosulfate leaching is applicable to the ores being persistent to the cyanide process: Manganese and copper-bearing ones. The optimal conditions shall be preserved maintaining pH-value at the level of 7-8 units in the solution. This ensures the stability of thiosulfate ions. In their absence, gold extraction sharply decreases, in addition, to increase the reaction rate it is recommended to enter elemental sulfur in the system.

Sorption on synthetic ion exchange materials is one of the most advanced methods for the hydrometallurgical extraction and concentration of noble and non-ferrous metals (Mohammadnejad *et al.*, 2011; Kononova *et al.*, 2014). For several decades a number of businesses have already been operating in the domestic industry using ion exchange resin sorption process both from the solutions and directly from ore slurries. The issues of ion exchanger regeneration and metal deposition from the resulting column effluents are addressed successfully. Experience gained has shown great prospects of the sorption technology.

Currently application of sorption method for noble and non-ferrous metals extraction from thiosulfate solutions is of utmost interest; this method allows increasing total extraction of noble metals from complex raw materials and prevents generation of toxic waste polluting the environment. Here the greatest difficulty is presented by selective metal extraction, which is complicated by the influence of impurity metals. In this regard, investigation into the selection of ion-exchange resins being able to selectively extract the target metals is an urgent task.

The range of sorbents showing capacity for ion exchange and donor-acceptor interactions and recommended for extraction of noble and non-ferrous metals from solutions is rather wide. However, absence of selective ion-exchange resins for gold extraction highly complicates the technology of its production. Investigated by the authors strong-base anion-exchange resins (AB-17, AM, IRA-400), having high sorption capacity and anion exchange ability in acidic and alkaline media, fail to meet requirements for selectivity of noble metal ions extraction. Quite a number of works are devoted to extraction of noble metals from chloride (Laatikainen and Paatero, 2005; Aktas and Morcali, 2011), thiocyanate (Li *et al.*, 2012; Kononova *et al.*, 2007) and thiosulfate (Fotoohi and Mercier, 2015) solutions using various sorbents.

Materials and Methods

While investigating sorption processes of non-ferrous (copper, iron, cobalt, zinc, nickel, manganese, lead) and noble (gold, silver) metals, chemical and electrochemical leaching solutions of concentrates from "A" and "M" deposits were used. Sodium thiosulfate solution was

applied as a solvent. A composite sulfur-graphite electrode was used as a source of leaching agent during preparation of electrochemical leaching solutions and it simultaneously ensured electrochemical reactions in the electrode-solution-mineral stock system (Kenzhaliev, 2014; Kenzhaliyev *et al.*, 2014). New anion-exchange resins were used as sorbents: AV-16 strong-base anionite, BNA-1 medium-base anionite, Purolite A500U/4994 and also AM-2B. Sodium thiosulfate having concentration of 0.2 M was used as a solvent. The experiments were carried out by stirring at a rate of 500 rpm, with anionite-to-solutions volume ratio of 1:500 and temperature 25°C.

Compositions of leaching solutions were determined using ContrAA 300 Flame Atomic Absorption Spectrometer. Spectra of leaching solutions in the ion form were obtained at Avatar 370 FT-IR Fourier Transform InfraRed Spectrometer in the spectral range 4000-400 cm^{-1} from the capillary layers in windows of KRS-5. Spectra were recorded from the layers with transmittance of 13-14% at a frequency of 1651 cm^{-1} .

Results

Sorption of Noble and Non-Ferrous Metals from Chemical Leaching Process Solutions

The findings of the investigations of metal sorption extraction from chemical leaching solutions of "A" deposit concentrate is shown in the Fig. 1.

As a result of studying metal sorption extraction process from leaching solutions of "M" deposit concentrate using sodium thiosulfate as a solvent and AV-16, BNA-1, AM-2B and Purolite A500U/4994 anionites (Fig. 2), it was revealed that the degree of Mn sorption by AB-16 anionite made 50% and sorption degree of other metals did not exceed 20% (Fig. 2a).

The degree of metal sorption extraction from the chemical leaching solution of "M" deposit concentrate on AM-2B (Fig. 2c) and Purolite A500U/4994 (Fig. 2d) ion-exchange resins is practically similar.

Sorption of Noble and Non-Ferrous Metals from Electrochemical Leaching Process Solutions

Degree of metal sorption extraction from electrochemical leaching solution of "A" deposit concentrate is shown in the Fig. 3.

During investigation of metals sorption behavior in the electrochemical leaching solutions of "A" deposit concentrate the best extraction results were obtained for zinc: 85% on AV-16 anionite (Fig. 3a), 99% on BNA-1 anionite (Fig. 3b) and practically 100% on Purolite A500U/4994 anionite (Fig. 3d).

Degree of metal sorption extraction from electrochemical leaching solution of "M" deposit concentrate with anionites is shown in the Fig. 4.

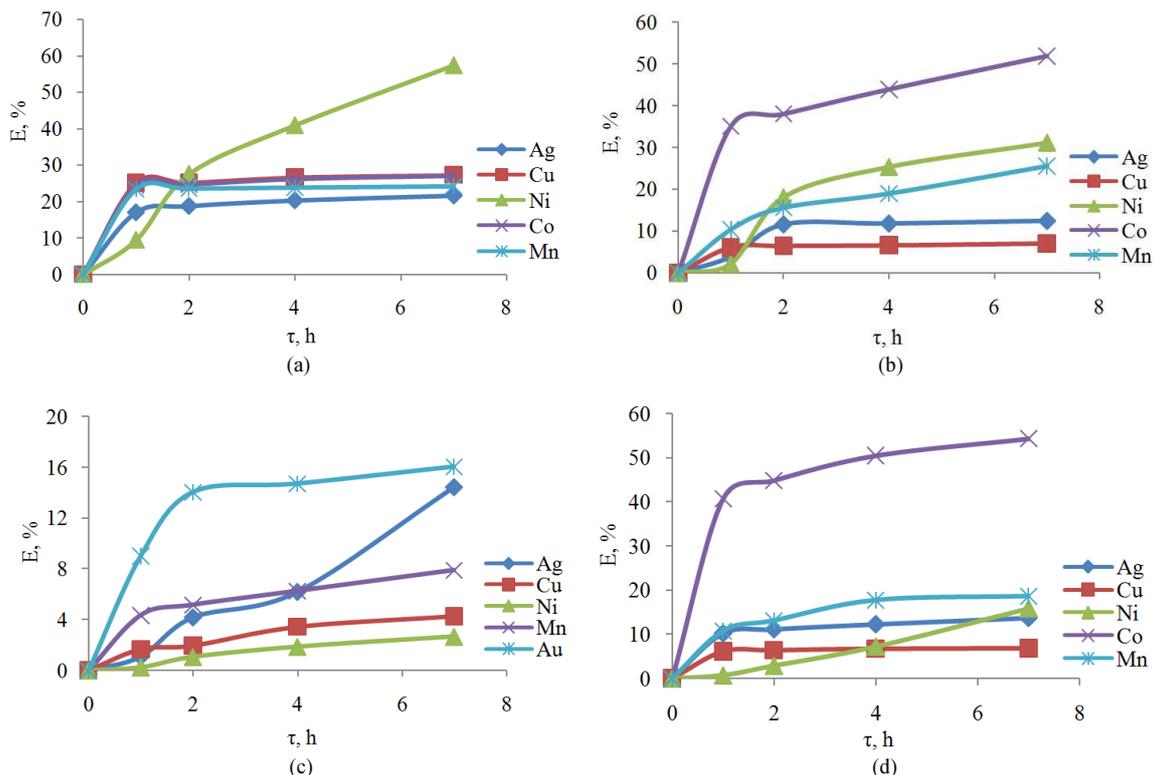


Fig. 1. (a) AV-16, (b) BNA-1, (c) AM-2B, (d) Purolite A500U/4994, degree of metal sorption extraction from the chemical leaching solution of "A" deposit concentrate with different ion-exchange resins

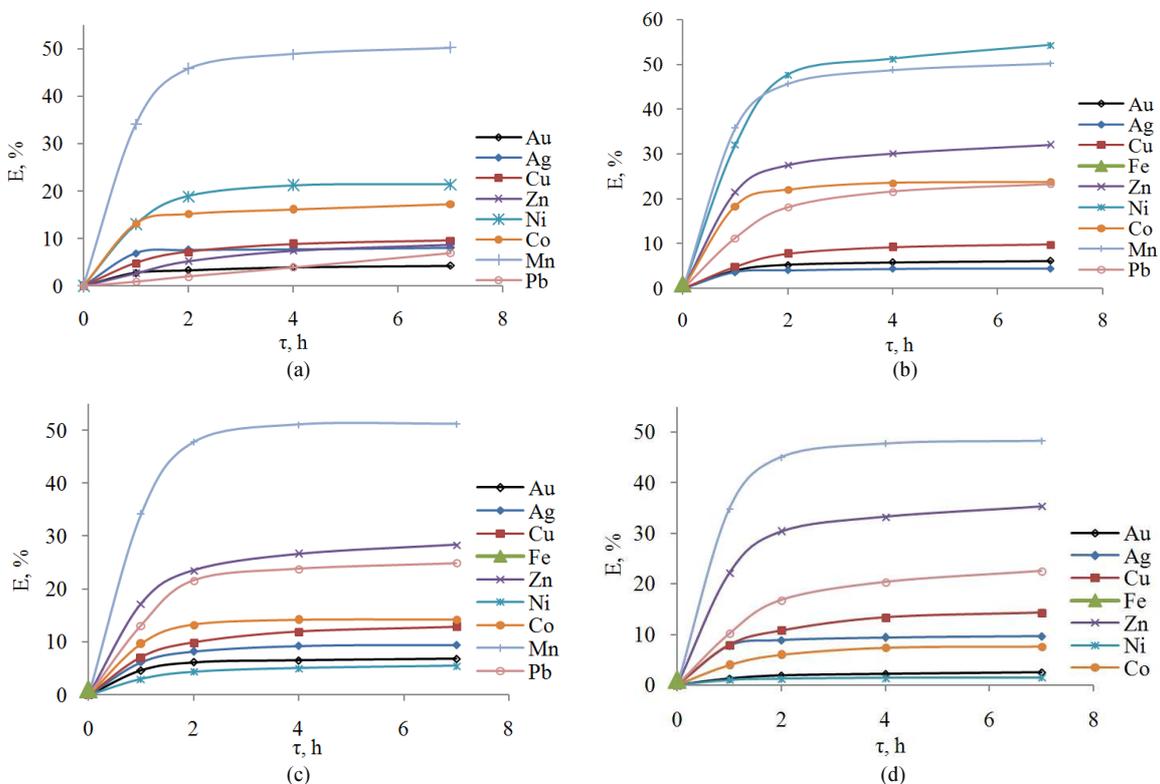


Fig. 2. (a) AV-16, (b) BNA-1, (c) AM-2B, (d) Purolite A500U/4994, degree of metal sorption extraction from the chemical leaching solution of "M" deposit concentrate

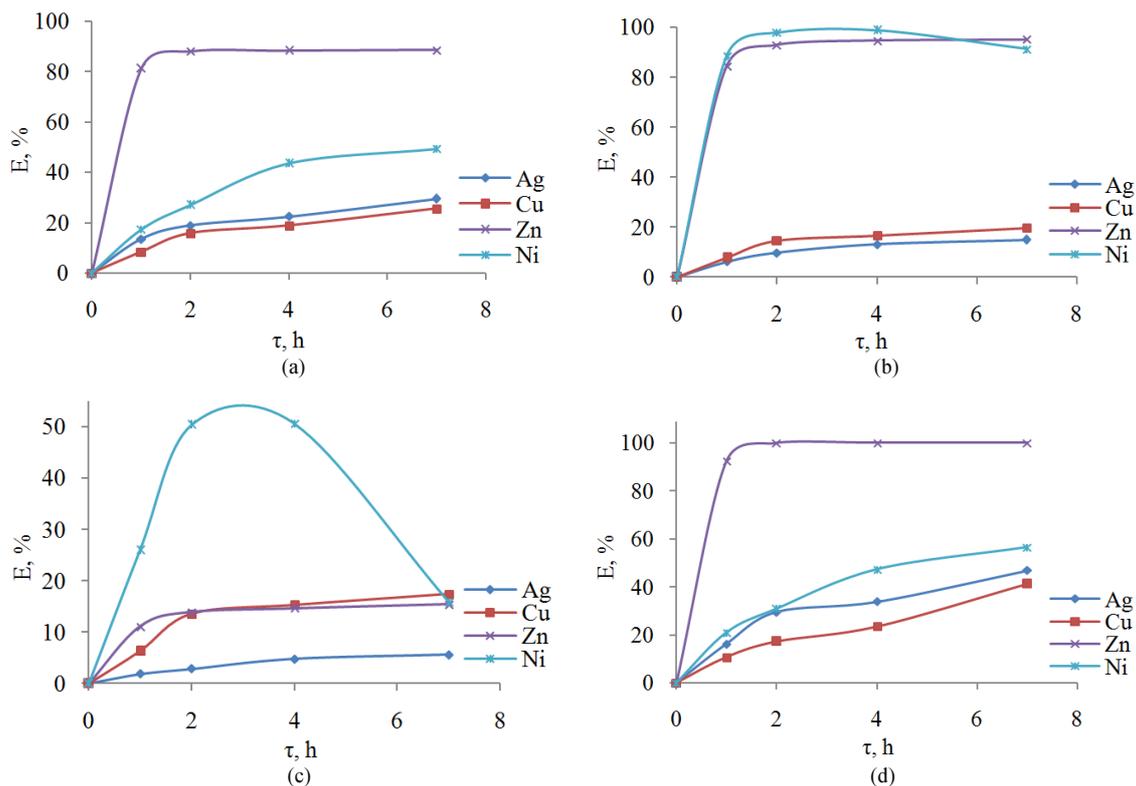


Fig. 3. (a) AV-16, (b) BNA-1, (c) AM-2B, (d) Purolite A500U/4994, degree of metal sorption extraction from electrochemical leaching solution of “A” deposit concentrate

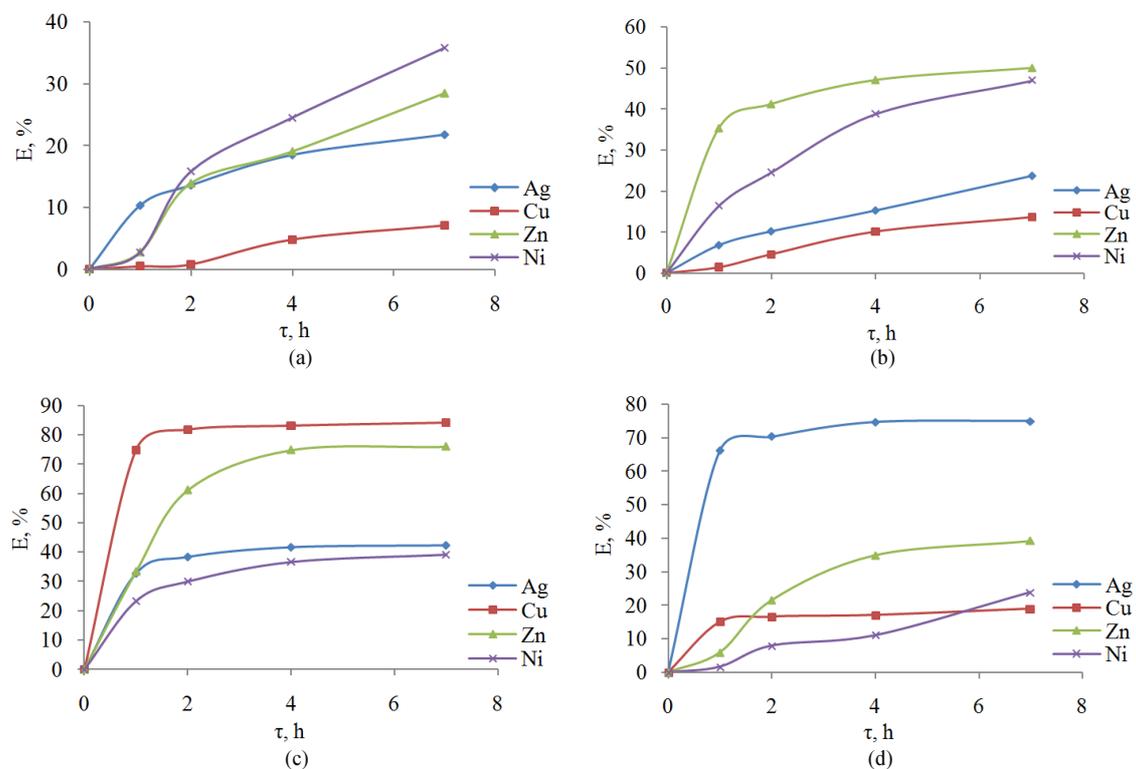


Fig. 4. (a) AV-16, (b) BNA-1, (c) AM-2B, (d) Purolite A500U/4994, degree of metal sorption extraction from electrochemical leaching solution of “M” deposit concentrate with anionites

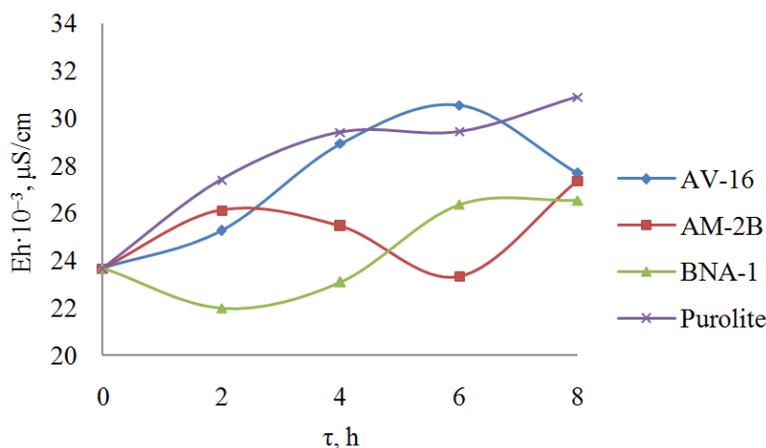


Fig. 5. Conductivity variation in the course of non-ferrous noble metals sorption from leaching process solution

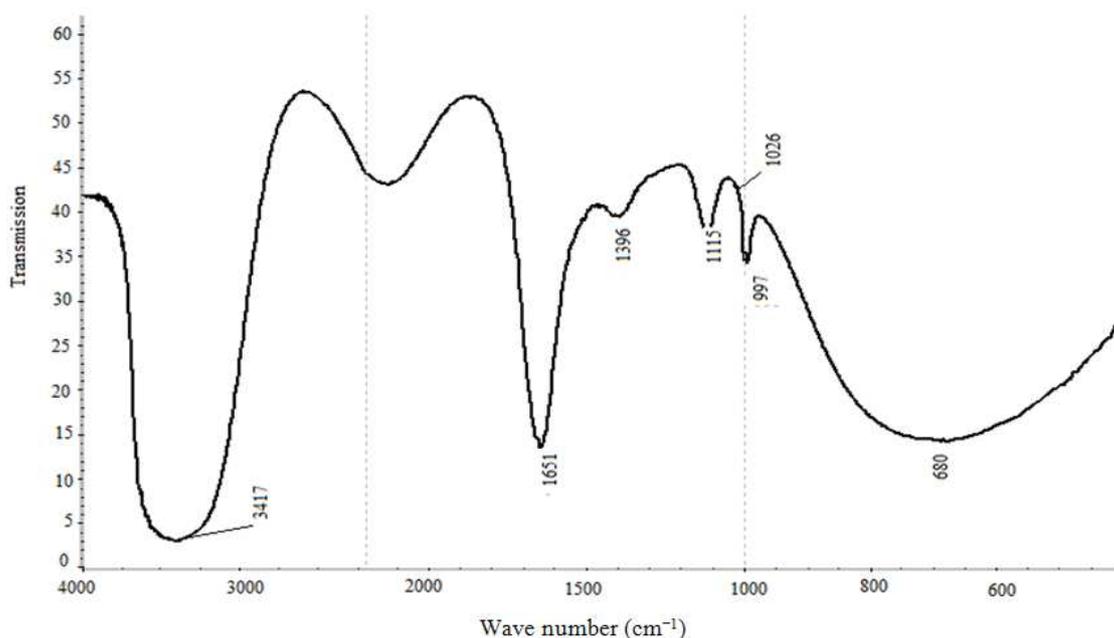


Fig. 6. IR-spectrum of electrochemical leaching process solution of “A” deposit concentrate Y-axis-Transmittance; X-axis-Wave number (cm^{-1})

In the course of non-ferrous and noble metals sorption electrical conductivity of solutions was periodically decreasing and increasing with continuous reduction of non-ferrous metals content in the solution (Fig. 5). This is connected both with sorption capacity of anion exchange resins and with variations in the microstructure of the investigated systems.

Compositions of Leaching Solutions

The results of atomic-absorption analysis of chemical and electrochemical leaching solutions of concentrates from “A” and “M” deposits are given in Table 1 and 2, respectively.

IR-Spectroscopy Research

R-spectrum of the electrochemical leaching process solution of “A” deposit concentrate is shown in Fig. 6. Absorption valence bands of $\nu(\text{OH})$, 3417 cm^{-1} , deformational bands of δHOH , 1651 cm^{-1} and librational band of $\nu_1\text{H}_2\text{O}$, 680 cm^{-1} of molecular water fluctuations were observed on the IR spectrum. Group $[\text{S}_2\text{O}_3]^{2-}$ was $1115, 997 \text{ cm}^{-1}$. Group $[\text{SO}_4]^{2-}$ was $1115, 1026 \text{ cm}^{-1}$ (Nakamoto, 2009).

Optical density was measured at the maximum of absorption bands that characterize fluctuations $\nu \text{ OH}$, $\delta \text{ HOH}$, $\nu_1\text{H}_2\text{O}$, $\nu_4(\text{E})$ of thiosulfate ion, $\nu_3(\text{F}_2)$ of sulfate ion, $\nu_1(\text{A}_1)$ of thiosulfate ion and ν_3 of $[\text{CO}_3]^{2-}$ group (Fig. 6).

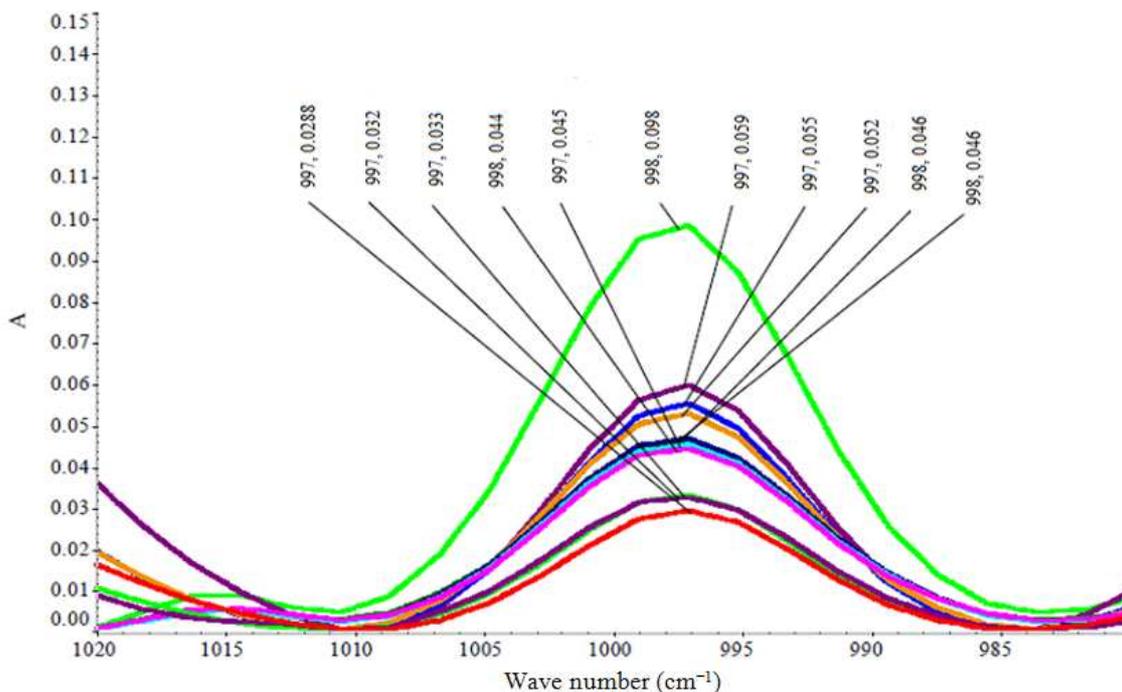


Fig. 7. Correlation between spectral regions of solutions after automatic adjustment of the baseline in the range of 1020-980 cm^{-1}

Table 1. Compositions of chemical leaching process solutions

| Solution | Au, mg/dm^3 | Ag, mg/dm^3 | Cu, mg/dm^3 | Ni, mg/dm^3 | Co, mg/dm^3 | Zn, mg/dm^3 | Mn, mg/dm^3 |
|---|-----------------------------|-----------------------------|-----------------------------|-----------------------------|-----------------------------|-----------------------------|-----------------------------|
| Chemical leaching solution of "A" deposit concentrate | - | 9.73 | 154.3 | 1.38 | 0.339 | - | 5.63 |
| Chemical leaching solution of "M" deposit concentrate | 0.0324 | 8.117 | 185.5 | 1.637 | 0.297 | 5.73 | 9.404 |

Table 2. Compositions of electrochemical leaching process solutions

| Solution | Ag, mg/dm^3 | Cu, mg/dm^3 | Zn, mg/dm^3 | Ni, mg/dm^3 |
|--|-----------------------------|-----------------------------|-----------------------------|-----------------------------|
| Electrochemical leaching solution of "A" deposit concentrate | 0.0480 | 11.86 | 0.039 | 0.124 |
| Electrochemical leaching solution of "M" deposit concentrate | 0.0613 | 78.80 | 0.530 | 0.490 |

Figure 7 shows correlation between spectral regions of chemical and electrochemical leaching process solutions before and after metal sorption by anionites upon automatic adjustment of the baseline in the range of 1020-980 cm^{-1} . The figure indicates spectra in the descending order of the intensity of absorption band of $[\text{S}_2\text{O}_3]^{2-}$ ion vibration $\nu_1(\text{A}_1)$. Intensity falloff is observed for $[\text{S}_2\text{O}_3]^{2-}$ ion vibration $\nu_1(\text{A}_1)$ absorption band from chemical and electrochemical leaching process solutions of concentrates to the solutions after sorption of non-ferrous and noble metals.

Discussion

In Fig. 1 illustrated that in case of using AV-16 anionite Ni is extracted to a greater extent (55%) (Fig. 1a), in case of BNA-1copper is primarily extracted (54%) (Fig. 1b) and Purolite A500U/4994 anionite

promoted cobalt extraction (59%) (Fig. 1d). AM-2B anionite demonstrated the least affinity to ions of metals contained in the chemical leaching solution of "A" deposit concentrate (Fig. 1c).

In case of BNA-1 anionite sorption degree of all metals increased during the first 2 h of leaching time and remained the same during next 6 h; in this case iron was not extracted and the extraction degree of other metals did not exceed 53% (Zn 27%, Ni 53%, Cu 8%, Co 23%, Zn 30%, Mn 48%, Pb 23%, Au 5% and Ag 4%) (Fig. 2b).

In both cases maximum extraction was characteristic of zinc (30 and 34%, respectively) and manganese (50%). Extraction of other metals on both anionites did not exceed 23% over 8 h of sorption.

Nickel was sorbed by 92% on anion exchange resins; BNA-1 anionite (Fig. 3b) showed Ni extraction degree (92%). In case of AM-2B anionite Ni extraction degree achieved 54% in the first 4 h of the

process, but after 4 h sorption degree fell down to 15% (Fig. 3c). This is likely to indicate resolution of nickel compounds.

In case of electrochemical leaching solutions of "M" deposit concentrate the degree of metal sorption extraction on AV-16 ion-exchange resin did not exceed 40% (Fig. 4a), on BNA-1 anionite metals were sorbed at most by 50% (Fig. 4b). For other anion-exchange resins these values were considerably higher; thus, in case of AM-2B copper was extracted by 84%, zinc-by 76% (Fig. 4c) and using A500U/4994 silver extraction degree of 75% was achieved (Fig. 4d).

Red shift of absorption band maximum value at wave number of 1115 cm^{-1} by 4 cm^{-1} relative to the fluctuations of $\nu_4(\text{E})$ of thiosulfate-ion in the 0.1 M test solution of sodium thiosulfate is indicative of the presence of $[\text{SO}_4]^{2-}$ ion in the system (Skoog *et al.*, 2014). The band with maximum at wave number of 1026 cm^{-1} in the range of manifestations of the stretching vibrations of S=O bond shows reduction of symmetry of sulfate group and monodentate coordination of this group [61, 62]. Group $[\text{CO}_3]^{2-}$ was 1396 cm^{-1} .

Optical density at the maximums of absorption bands assigned to water stretching vibrations $\nu\text{ OH}$ amounted to 1.541; OD assigned to water deformation vibrations $\delta\text{ HOH}$ equaled to 0.871; OD assigned to water libration vibrations $\nu_{\text{L}}\text{H}_2\text{O}$ was 0.850. At the maximum of absorption band characterizing vibration $\nu_4(\text{E})$ of thiosulfate ion, at wave number of 1115 cm^{-1} optical density made 0.424^1 (Nakamoto, 2009).

Intensity falloff is observed for $[\text{S}_2\text{O}_3]^{2-}$ ion vibration $\nu_1(\text{A}_1)$ absorption band from chemical and electrochemical leaching process solutions of concentrates to the solutions after sorption of non-ferrous and noble metals (Fig. 7).

Conclusion

This paper presents a study that has been undertaken to investigate sorption behavior of noble and non-ferrous metals from chemical and electrochemical leaching thiosulfate solutions using various anionites. The potential selective sorption extraction of zinc and nickel from electrochemical leaching solutions on BNA-1 anionite was revealed, which will enable to obtain practically complete extraction thereof (up to 99%) and will contribute to increasing the extraction degree of other valuable components (copper, silver) resulting from concentration reduction of metals, having high affinity to anionite.

Percentage of gold and silver extraction from thiosulfate solutions is much less than the same for cyanide solutions, which is attributable to the specific sorption conditions: High concentration of thiosulfate ions (as well as to the presence of S^{2-} , SO_4^{2-} , $\text{S}_4\text{O}_6^{2-}$ and

SO_3^{2-} ions formed as a result of thiosulfate decomposition) and presence of thiosulfate complexes of non-ferrous metals in the system.

It has been demonstrated that depending on the type of the raw material used it is efficient to perform silver and copper sorption from electrochemical leaching solutions with the help of Purolite A500U/4994 and AM-2B anion-exchange resins. In case of chemical leaching process solutions, the degree of metal extraction is less, which is connected with more complex chemical composition of these solutions.

It is obvious that the nature of the sorption extraction of metals from process solutions is associated with the microstructure of the solutions, since leaching solutions have quasi-homogeneous structure. This fact is confirmed by the extreme variation of solution conductivity with continuous reduction of metals content in the solution.

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Author's Contributions

Bagdaulet Kenzalyevich Kenzhaliyev: Contributed to the reviewing of the article critically.

Renata Ramilevna Iskhakova: Organized the study, data collection of the study sample, analysis and writing of the manuscript.

Zamzagul Duisenovna Dosymbayeva: Designed the research plan, participated in the results analysis.

Ethics

The authors have no conflicts of interest in the development of the research and publication of this article.

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